SOLVENT TOXICOLOGY: RECENT ADVANCES IN THE TOXICOLOGY OF BENZENE, THE GLYCOL ETHERS, AND CARBON TETRACHLORIDE

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INTRODUCTION

The field of solvent toxicology is quite broad and impossible to review fully in this format. We decided to stress benzene, the glycol ethers, and carbon tetrachloride. This review emphasizes the projected role of xenobiotic metabolism in hemopoietic toxicity caused by benzene, adverse effects on the male reproductive system caused by the glycol ethers, and hepatotoxicity caused by carbon tetrachloride.

BENZENE METABOLISM AND TOXICITY

Introduction

Benzene (Bz) is an important industrial chemical, a petroleum by-product, a component of unleaded gas, and thus a ubiquitous environmental pollutant [reviewed in (1)]. Bz is a myelotoxin; chronic exposure of humans and experimental animals to high concentrations results in blood dyscrasias including lymphocytopenia, thrombocytopenia, and pancytopenia or aplastic anemia (2, 3). Bz is also a carcinogen. It is associated with an increased incidence of acute myelogenous leukemia and some of its variants in humans (4–10), an increased incidence of several solid tumors (11, 12), and possibly leukemia/lymphoma (13) in rodents.

Since recent articles have reviewed the historical aspects of benzene toxicity (14), its genotoxicity (15), quinones as its toxic metabolites (16), and its developmental toxicity (17), we review recent findings on the metabolism of benzene and its hematotoxicity.

Metabolism

Benzene metabolism is required for toxicity. Possible pathways for the bioactivation of Bz in vivo are shown in Figure 1. In the liver, the major site of Bz metabolism (18), Bz is converted via a cytochrome P-450-mediated pathway (19) to benzene oxide, which is transformed by epoxide hydratase to the 1,2-dihydrodiol, which leads to catechol (C) formation (20) or rearranges nonenzymically to phenol (P), which is metabolized to hydroquinone (HQ) (21). The bone marrow is the target organ. It possesses a limited capacity to metabolize Bz, which cannot account for the amount of metabolites that accumulate in it (22–24). P is metabolized to HQ and C in the marrow by a myeloperoxidase-mediated pathway (25). The oxidation of HQ to p-benzoquinone (BQ) also probably occurs by a myeloperoxidase-mediated pathway, whereas C is presumably converted to 1,2,4-benzenetriol (BT) by the cytochrome P-450 system.

The quinones or semiquinones derived from HQ and/or C are generally considered to be the toxic metabolites, although an open-ring product such as *trans,trans*-mucondialdehyde is a possible toxic metabolite (26, 27). Ethanol consumption potentiates Bz toxicity in vivo by accelerating the hydroxylation of Bz and the conversion of P into toxic metabolites (28). Thus, it increases the hematotoxicity.

MICROSOMAL METABOLISM OF BENZENE Bz induces the enzymes required for its metabolism. Post & Snyder (29) observed two benzene hydroxylase activities in rat liver microsomes from control, beta-naphthoflavone (BNF)-treated, and Bz-treated animals at all Bz concentrations and in pheno-

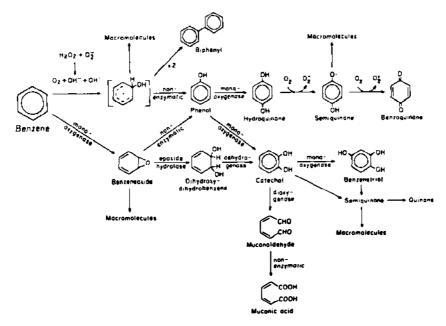


Figure 1 Metabolic pathways of benzene (taken from Ref. 45 with permission).

barbital (PB)-induced microsomes at concentrations of Bz up to 0.8 mM. Bz pretreatment increased activity without affecting total cytochrome P-450 content. This finding suggests the induction of a specific isozyme that was saturated at Bz concentrations > 0.4 mM, had a pH optimum of 6.6, and was stimulated by fluoride (30).

The inductive pattern of Bz hydroxylase is influenced by the addition of methyl groups to Bz (31). Toluene and xylene increased the metabolism of Bz, and consistent with the report by Post & Snyder (29), Bz induced benzene hydroxylase, but did not induce aniline hydroxylase or aminopyrine *N*-demethylase, whereas toluene and xylene did. Bz more readily induced the conjugating enzymes and increased the level of glutathione (GSH).

Cytochrome P-450 fractions from microsomes of benzene-induced rats have been compared to those obtained from untreated or PB-, BNF-, methyl-cholanthrene-, and ethanol-induced animals (32). Four distinct cytochromes P-450 were identified; isozyme A was induced by Bz and ethanol, and isozyme Bb was induced by all the inducers. The Bb isozyme induced by Bz appeared to differ from that induced by the classical inducers. The molecular activity of benzene hydroxylase of the microsomal fraction of rabbit bone marrow was about four times higher than that of liver microsomes (33). Pretreatment of the animals with Bz, but not PB, induced benzene mono-oxygenase activity, but did not affect the O-dealkylation activity in the marrow. Covalent binding of ¹⁴C-labeled metabolites, but no free HQ or C,

could be detected in incubations of Bz with bone marrow microsomes from Bz-induced animals (33).

Gilmour et al (34) found that Bz was converted to P by rat liver microsomes. They also observed that small amounts of HQ and C were produced from both Bz and P in a reaction mediated by cytochrome P-450 and stimulated by Bz pretreatment. The metabolism of Bz and P was competitively inhibited by toluene, and Bz and P reciprocally inhibited the metabolism of each other. Lunte & Kissinger (35) found that the rate of conversion of P to HQ by mouse liver microsomes was significantly higher than the rate of metabolism of Bz. P was converted to HQ, which was oxidized to p-BQ. Benzene hydroxylation in rabbit liver microsomes and reconstituted membrane vesicles containing cytochrome P-450 LM2 appeared to involve hydrogen peroxide, superoxide anion, and hydroxyl radical (36). Biphenyl was formed by the reconstituted system, indicating a cytochrome P-450-dependent formation of a hydroxycyclohexadienyl radical from the interaction of hydroxyl radical with Bz. These results would suggest that the microsomal cytochrome P-450-dependent oxidation of Bz requires hydroxyl radicals generated in a modified Haber-Weiss reaction between hydrogen peroxide and superoxide anion. However, these studies used only 17- μ M Bz. Using a similar system with higher concentrations of Bz in alcohol, Gorsky & Coon (37) found a $K_{\rm m}$ for Bz of 18 mM with PB-induced rabbit liver microsomes and 105 mM with a cytochrome P-450 LM2 reconstituted system. When Bz concentrations were in the range of the $K_{\rm m}$, superoxide dismutase, desferrioximine, or hydroxyl radical scavengers had no effect. The oxidation of Bz to P in a model hydroxyl radical-generating system consisting of xanthine, xanthine oxidase, and Fe-EDTA was not dependent on the substrate concentration. The rate of hydroxyl radical generation by the model system was regulated to be greater than the rate of product formation in the microsomal systems. Therefore, the lack of dependence on Bz concentration suggests that hydroxyl radicals are not involved in the metabolism of Bz to P when the concentration of Bz is near the $K_{\rm m}$ for cytochrome P-450 LM2. At concentrations of Bz below the $K_{\rm m}$, the free-radical pathway of P formation becomes increasingly predominant.

It is commonly assumed that the microsomal cytochrome P-450 conversion of Bz to P occurs via an epoxide intermediate (38, 39), demonstrable by an NIH shift. Hinson et al (40) studied the role of the NIH shift in the formation of P from the deuterated benzene derivative, 1,3,5-[²H₃]benzene, and found the expected products, 2,3,5-[²H₃]phenol and 2,4-[²H₂]phenol, which indicated that the shift had occurred. Analysis of the deuterium-isotope effect in the deuterated phenols suggests that the cyclohexadienone was formed either

by the somerization of the epoxide or directly from the enzyme-substrate complex as a major intermediate in the hydroxylation of Bz.

Both epoxidation and free-radical insertion may be involved in the formation of P from Bz. Griffiths et al (41) reported on the production of P from Bz by a reconstituted microsomal system containing purified rat liver PB-induced cytochrome P-450. This production was inhibited by metyrapone and SKF 525A, but the conversion of P to secondary metabolites was not. The formation of the polyhydroxymetabolites from P was inhibited by radical-trapping agents and desferrioxamine. Interaction of Bz with cytochrome P-450 in the reconstituted microsomal system elicited a Type I spectral change, whereas interaction of P produced a Type II change indicative of an interaction of the substrate with the iron moiety. Also, an excess of unlabeled P did not dilute the labeled P derived from labeled Bz. These facts suggest that Bz interacts at the active site of cytochrome P-450 and undergoes epoxidation to benzene oxide, which rearranges to form P. This P is preferentially retained on the cytochrome P-450 but is shifted from the catalytic site to the iron moiety. There it can be attacked by hydroxyl radicals to form the polyhydroxy metabolites.

MITOCHONDRIAL METABOLISM OF BENZENE Rat liver mitochondria, stripped of their outer membrane to avoid microsomal contamination (mitoplasts), metabolize Bz in an NADPH-dependent reaction to P and metabolites that covalently bind to mitochondrial (mt)DNA (42). A cytochrome P-450 that converts Bz to P has been solubilized from mitoplasts with 0.4% sodium cholate and purified 23-fold by polyethylene glycol (PEG) fractionation (43). The production of P from Bz by the 5–15% PEG fraction is completely dependent on NADPH and an exogenous bovine adrenodoxin/adrenodoxin reductase system that microsomal cytochrome P-450s do not use. This dependence provides evidence that the activity is indeed mediated by a cytochrome P-450 of mitochondrial origin.

THE EFFECT OF BENZENE AND ITS METABOLITES ON REPLICATION AND Bz and its metabolites inhibit both nuclear and mitochondrial replication and transcription. DNA synthesis was inhibited in hemopoietic cells from mice exposed to a single dose of 3000-ppm Bz (44). It was also inhibited in mouse L5178YS lymphoma cells after their exposure to the metabolites, but not to Bz, which is not bioactivated in these cells (45). BQ was the most potent inhibitor, followed by HQ, BT, C, and P at concentrations that were not cytotoxic. Inhibition correlated with ease of oxidation. This correlation suggests that the oxidation of P or one of its metabolites produces the ultimate reactive compound that inhibits DNA

synthesis. Only BQ and BT induced single-strand (ss) DNA damage (46). BQ was about nine times more effective than BT in causing the damage. Ascorbic acid (47) protected against damage from BQ; superoxide dismutase did not protect against damage from BQ, but decreased the DNA damage from BT. Inhibition of intracellular superoxide dismutase increased the ssDNA damage. Such an increase suggests that ssDNA damage caused by BT results from superoxide anion radical generated during BT oxidation. Glutathione (GSH) also protected against DNA damage (48); BQ rapidly depleted the intracellular store of GSH. During BT treatment the GSH level dropped gradually, and only BQ affected the level of oxidized glutathione (GSSG). These observations imply that the two compounds act on GSH through different mechanisms and that GSH plays an important role in the detoxication mechanism (or mechanisms) for BQ and BT.

DNA replication in rat liver and rabbit bone marrow mitochondria in vitro is inhibited in a dose-dependent manner ky HQ, p-BQ, and BT (49). The activity of purified rat liver mitochondrial DNA polymerase- γ was inhibited by HQ and BQ by the metabolite's interaction with an active sulfhydryl group on the enzyme. The binding of [14 C]hydroquinone to the enzyme was prevented by N-ethyl maleimide as well as by unlabeled HQ or p-BQ, suggesting that both compounds bind to the same sulfhydryl (SH) group on the polymerase or that they are interconverted to the benzosemiquinone, which then binds.

Transcription in mouse lymphocytes (50) and macrophages (51) in vitro is inhibited in a dose-dependent manner by P, HQ, and p-BQ; the IC_{50} for BQ was 5×10^{-6} M for both cell types. Translation was also inhibited subsequent to the inhibition of RNA synthesis. Phenol was metabolized in macrophages by peroxidase to reactive species that inhibited RNA synthesis and covalently bound to macromolecules (51). Benzene was not metabolized. It inhibited RNA synthesis by infiltrating the plasma membrane and preventing the transport of labeled uridine into the cell, in a manner similar to that observed for lindane (52).

Rat and rabbit liver and cat and rabbit bone marrow mitoplasts are capable of bioactivating Bz in vitro to metabolites that inhibit mitochondrial transcription (53) and, consequently, translation. HQ, BQ, or P also caused a dose-dependent inhibition of transcription in rabbit bone marrow mitoplasts, with IC_{50} values ranging from 2×10^{-6} M for p-BQ to 2×10^{-3} M for C. Similar concentrations of p-BQ, HQ, and C inhibited mRNA synthesis in rabbit bone marrow nuclei incubated under conditions specific for RNA polymerase II (54).

COVALENT BINDING OF REACTIVE METABOLITES OF BENZENE WITH MAC-ROMOLECULES Bz and certain Bz metabolites are converted in nuclei and mitochondria to reactive species that covalently bind to macromolecules and that cause DNA damage in the form of strand breaks and covalently bound adducts in liver (55) and bone marrow (56). Reactive species from [14C]benzene administered to rats and mice covalently bind to the macromolecules of various organs (57). Binding to RNA and protein is an order of magnitude greater than that to DNA. The level of DNA binding was low in several organs and highest in bone marrow. The covalent binding index (CBI) (cf 55) was 10, which approximately equals that of genotoxic carcinogens classed as weak initiators of carcinogenesis. The microsomal metabolism of Bz to reactive species capable of covalently binding to nucleic acids and proteins was induced by PB and inhibited by SKF 525A and GSH. This metabolism was selective. It was mediated by microsomes from liver but not from kidney, spleen, or lung (57). Ascorbate significantly reduced the ability of PB- and Bz-induced hepatic microsomes to metabolize labeled P or Bz to species capable of covalently binding to microsomal proteins (58). The metabolism of Bz to P was unaffected by GSH. However, GSH inhibited the covalent binding species observed from both substrates more than 95%. The metabolism of P to HQ was unaffected by either GSH or ascorbate. Bone marrow from guinea pigs with low dietary intakes of ascorbate showed a fourfold greater covalent binding of phenol equivalents in the presence of peroxide than did marrow from animals with high intakes. Myeloperoxidase appears to be responsible for the oxidation of P, as ascorbate blocks the oxidation of P, and H₂O₂ is for required the activation of P to covalently binding species. Partially purified guinea pig DT-diaphorase [NAD(P)H quinone oxidoreductase] inhibited the covalent binding from P by 70%. This effect was reversed by the diaphorase inhibitor, dicumeral (59). On the other hand, covalent binding was greatly enhanced by H2O2 and horseradish peroxidase or myeloperoxidase, and prevented by ascorbate. Quinone oxidoreductases reduce quinones to HQ or C via a two-electron transfer. This reduction suggests that covalent binding results from an electrophilic quinone species rather than from a semiquinone and the selective myelotoxicity of Bz may result from the high ratio of peroxidase activity to quinone reductase activity in the bone marrow (59). Lack of Bz toxicity in the liver may thus result from the conversion of quinone metabolites back to HQ and C by DT-diaphorase or by carbonyl reductase. Carbonyl reductase is very active on p-BO and is the major NADPH quinone reductase of human liver (60).

Mitochondria can also activate Bz to intermediates that bind covalently to mtDNA. Rat liver mtDNA contains six deoxyguanosine adducts, and rabbit bone marrow mtDNA contains seven such adducts after incubation of mitoplasts with labeled Bz (61). A hydroxyl radical scavenger, mannitol, prevented the formation of four deoxyguanosine adducts. Two deoxyadenosine adducts were formed from [¹⁴C]benzene in vitro. HQ, p-BQ, C, and Bz also

formed adducts with deoxyguanosine when incubated with bone marrow mitoplasts in vitro (61).

Jowa et al (62) characterized the adducts formed when HQ or p-BQ reacts with deoxyguanosine (dG) in vitro. Two dG-adducts were formed with both HQ and p-BQ. Iron was required for the formation of the the adducts from HQ but not from p-BQ. This difference indicates the necessity for oxidation of HQ to BQ or the semiquinone. Similar adducts of dG were formed when p-BQ was allowed to react with single- or double-stranded DNA or with chromatin from rat liver nuclei. Adduct 2 was more stable. It was further characterized by NMR and mass spectra analysis to have the structure presented as Structure 1.

HEMOPOIETIC TOXICITY Aplastic anemia from Bz poisoning could arise from toxic damage to one or more of the components of the hemopoietic system: stem cells, transit cells (progenitor cells in various degrees of differentiation), and/or bone marrow stroma or microenvironment (63).

Effects on stem cells Several investigators have reported reductions in the number of multipotential stem cells (CFU-S) following exposure of mice to Bz (64, 65), whereas other workers have reported no change (66, 67). This discrepancy has been explained, in part, by a toxic effect of Bz on the transit cells, which signals stem-cell proliferation and differentiation, and thus depletes the stem-cell pool (68–70).

Effects on progenitor cells Evidence that Bz acts primarily on committed progenitor cells was indicated by a decrease in the number of differentiating erythroid and myeloid progenitors in benzene-exposed rats, without an effect on the number of progenitor cells or mature cells in the bone marrow (71). Studies on the effects of Bz on the kinetics of ⁵⁹Fe uptake into maturing mouse erythrocytes indicated that the pronormoblast was the progenitor most sensitive to benzene, whereas stem cells and nondividing reticulocytes were relatively unaffected (72). The metabolites P, C, and HQ have also been reported to reduce ⁵⁹Fe incorporation into developing erythrocytes significantly, but not as much as Bz. The mixed-function oxidase inhibitor, 3-amino-1,2,4-triazole abolished the erythropoietic toxicity of Bz and P, but not of C or HQ (73). *Trans,trans*-mucondialdehyde, a six-carbon alpha, beta unsaturated aldehyde and an open-ring metabolite of Bz, is hemotoxic in CD-1 mice in a manner similar to benzene (26). It is also cytotoxic to human

erythrocyte progenitor cells at micromolar concentrations (27). Other studies have shown that the number of erythroid (CFU-E) (74) and granulocyte/macrophage (CFU-GM) (75) progenitor cells was depressed in benzene-exposed mice. However, low-level (10-ppm) exposure over a 54-day period did not decrease the number of CFU-E or CFU-GM unless hemopoiesis was stimulated by administration of a hemolytic dose of phenylhydrazine (76).

Effects on the stromal microenvironment Hemopoiesis results from the interaction of stem cells with the supporting stroma. The stroma provides a favorable microenvironment for the regulated proliferation and differentiation of the stem cells (77). Reconstitution of the stroma and stem cells in vitro (78, 79), based on the growth of an adherent stromal layer from bone marrow cells in liquid culture, has stimulated new experimental approaches and provided important data about benzene toxicity to the bone marrow.

Several reports have indicated that Bz is toxic to the marrow microenvironment. Injection of normal bone marrow cells into benzene-exposed, lethally irradiated mice could not reconstitute normal hemopoiesis, as the number of granulocyte and macrophage colonies produced in vivo was decreased (66). However, incubation of normal bone marrow cells with Bz prior to administration to the animals did not reduce colony formation. This finding suggests that the microenvironment was deranged by benzene. Garnett et al (80) demonstrated that the marrow-adherent layer from benzene-treated mice did not differ in the number of CFU or in their ability to proliferate but was less able to support the differentiation of stem cells from unexposed mice. The development of an adherent layer in culture from the marrow cells of Bztreated animals was altered; fat cells that normally appear during growth fail to develop, indicating that Bz affects marrow stromal cells. Gaido & Wierda (81) showed that HQ and p-BQ were most effective in decreasing the ability of stromal cells to support CFU-GM colony formation in a coculture system, whereas C and BT only inhibited at high concentrations. Adherent stromal cells from relatively Bz-resistant B6C3F1 mice support hemopoiesis better than stromal cells from sensitive DBA/2J mice (82). Marrow cellularity was reduced more in D2 than in B6 mice, but Bz had no effect on adherent stromal cell colonies or the number of granulocyte/macrophage precursors present. P, but not Bz, significantly decreased the ability of adherent stromal cells to support hemopoiesis of granulocyte and macrophage precursors, but no strain difference was evident. HQ, at levels that did not alter stromal cell number, inhibited granulocyte/macrophage colony formation and increased prostaglandin E₂ (PGE₂) levels (83). Pretreatment of the cultures with indomethacin decreased PGE₂ levels and protected against toxicity. Thus, the authors (84) suggested that HQ suppression of stromal cell-supported hemopoiesis is mediated by increased PGE₂ production and that this increase may be involved in myelosuppression by Bz. Taken together, these studies indicate that injury to bone marrow stromal cells may be an important factor in benzene-induced myelosuppression.

A possible target of Bz toxicity in the stroma is the macrophage. Macrophages are a major source of polypeptide growth factors required for the proliferation, development, and survival of progenitor cells of the various hemopoietic lineages (85). Post et al (51) demonstrated that P is metabolized in adherent mouse marrow macrophages by a peroxidase activity to one or more covalently binding species and that micromolar concentrations of HQ and p-BQ inhibit macrophage RNA synthesis. Adherent monocyte and macrophage cells from benzene-treated rabbits have been reported to inhibit the development of CFU-E and BFU-E colonies when cocultured with normal bone marrow (86). This finding suggests that Bz induces a CFU-E and BFU-E inhibiting activity in adherent blood cells.

Effect of benzene on lymphopoiesis Peripheral lymphocytopenia is an early manifestation of Bz toxicity in both animals and humans (6, 16, 87), and is a distinctive feature of benzene-induced aplastic anemia (3). The ability of P, HQ, and C to suppress lymphocyte growth and function in vitro correlates with their capacity to undergo autooxidation and with their concentration in the bone marrow or lymphoid organs (88–91). HQ and its oxidation product, p-BQ, inhibit proliferation and differentiation in lectin-stimulated lymphocytes in culture (89, 90). These compounds also interfere with microtubule assembly (88, 91) at concentrations that are not cytotoxic and that can be achieved in vivo. P and C suppress lymphocyte activation only at cytotoxic concentrations. Suppression of lymphocyte blastogenesis by HQ has been postulated (88) to be mediated by the interaction of p-BQ with sulfhydryl groups on tubulin. This binding interferes with microtubular integrity, which is essential in cell division via spindle formation and in the regulation of surface receptor movement and signal transduction across the plasma membrane.

HQ and C are also immunotoxic in vivo. They are cytotoxic to spleen cells (92). HQ and C inhibit dextran sulfate— or lipopolysaccharide (LPS)-induced development of polyclonal plaque-forming cells (PC-PFC). These PC-PFC were obtained from progenitor cells from spleen and marrow of benzene-treated animals that were induced to differentiate by dextran sulfate or LPS. Both metabolites reduced the number of PC-PFC that developed from progenitors obtained from these organs. However, only C inhibited LPS-activated marrow progenitors from maturing to PC-PFC. HQ also inhibited pre-B cells (IgM⁻) from maturing to (IgM⁺) cells, and HQ reduced the ability of mitogens to stimulate the proliferation of IgM⁺ cells to CFU-B colonies (93). Thus, HQ and C, by a reduction of progenitor B-lymphocytes, are immunotoxic in vivo, and HQ inhibits marrow lymphopoiesis in vitro at a

concentration of 10^{-7} M. This concentration can be achieved in vivo, suggesting that inhibition of precursor-cell maturation is important in the hemotoxicity from Bz (92). Proliferation and maturation of lymphocyte progenitor cells are regulated by polypeptide lymphokines, which are produced both in vivo and in vitro by T-lymphocytes. Bz, if metabolized in the lymphocyte to a reactive intermediate such as p-BQ, could inhibit the production of lymphokines. Post et al (50) demonstrated that HQ and p-BQ affected the dose–dependent inhibition of RNA synthesis in mouse spleen lymphocytes in vitro at micromolar concentrations that were not cytotoxic. They also demonstrated that exposure to p-BQ completely inhibits the proliferation and production of the T-cell lymphokine, interleukin-2, by concanavalin (conA)-stimulated T-lymphocytes.

A number of in vivo studies have been conducted by C. A. Snyder and his collaborators on the ability of Bz to affect lymphopoiesis and modulate the cell-mediated immune response. Protracted exposure of mice to a regimen known to cause thymic lymphomas decreased B-lymphocytes in bone marrow and spleen. It also decreased T-cells in thymus and spleen and the ability of T-cells to respond to mitogenic stimulation (94). Bone marrow cellularity was increased 3-fold and the numbers of thymic T-cells, 15-fold, whereas no compensatory response was seen in the spleen. This difference led to the speculation that a subpopulation of thymocytes may exist that are resistant to Bz and that proliferate in its presence (94). More importantly, short-term exposure of mice to Bz concentrations at or near the industrial standard-exposure level significantly depressed mitogen-induced blastogenesis of both B- and T-lymphocytes without reducing the total number of either type of cell (95).

Bz can modify both host resistance to a bacterial infectious agent (96) and T-cell-mediated tumor resistance (97). The immunosuppressive effects of Bz have been reported to be modulated by the prior administration of a fungal product, 6FMA, from Aspergillus ochraceous that has interferon-inducing properties (98). Ingestion of ethanol also increases immunosuppresion by Bz (28), as well as benzene-induced hematotoxicity (99) in experimental animals. If these findings can be extrapolated to humans, they might have considerable impact on workers who are exposed to Bz on a daily basis and who are moderate-to-heavy drinkers.

METABOLISM AND TOXICITY OF GLYCOL ETHERS

Introduction

The glycol ethers (GE; Table 1), often referred to as cellosolves, are an important class of industrial solvents that are miscible with both water and many organic solvents (100). Because of their low vapor pressure and high

Name	Abbreviation	Synonyms	Structure
Ethylene glycol monomethyl ether	EGME	2-methoxyethanol Methyl Cellosolve®	СН₃ОСН₂СН;
Ethylene glycol monoethyl ether	EGEE	2-ethoxyethanol Cellosolve®	CH ₁ CH ₂ OCH ₂ C
Ethylene glycol monobutyl ether	EGBE	2-butoxyethanol Butyl Cellosolve®	CH ₃ (CH ₂) ₃ OCH ₂ ⁰
Propylene glycol monomethyl ether	PGME	I-methoxy-2- propanol	СН₃ОСН₂СНО

Table 1 Names and structures of some important glycol ethers

rate of dermal absorption, significant exposure can occur through contact with the skin (101).

The older literature (101) and genotoxicity (102) of the GE have been reviewed recently. The genotoxic potential of these compounds, if any, is minimal. Occupational exposure to ethylene glycol methyl ether (EGME) has been associated with CNS, hematopoietic, and renal toxicity; few cases of illness from other GE have been reported.

Recently, attention has focused on the effects of the GE on the male reproductive system, on fetal and embryonic development, and on the hematopoietic system. We review these areas. Although testicular toxicity was first noted fifty years ago (see 103), the potential of the GE to cause this toxicity was not widely appreciated until much later. Reproductive effects were reviewed in 1983 (103); since then, much new information has become available.

Male Reproductive Effects

resticular toxicity of EGME and EGEE (104–115) are summarized in Table 2. EGME and ethylene glycol ethyl ether (EGEE) cause testicular atrophy, decreased sperm counts, abnormal sperm motility and morphology, degeneration and atrophy of the seminiferous tubules, and impairment of fertility. EGME is more potent than EGEE (104, 107, 116). Their acetate derivatives are as potent testicular toxins as the parent compounds, presumably because the ether linkage is readily hydrolyzed (104). Other GE including ethylene glycol monobutyl ether (EGBE) (104, 111, 117–119), ethylene glycol phenyl ether (104), ethylene glycol isopropyl ether (112), ethylene glycol mono-*n*-propyl ether (120), propylene glycol monobutyl ether (PGME) (105, 111), and dipropylene glycol monomethyl ether (DPGME) (121) did not cause testicular damage.

Table 2	Testicular	toxicity.	caused 1	bν Ε	GME	and	EGEE ^a

Compound	Species	Exposure	Dose ⁶	Reference
EGME	mouse	oral, 5 d/wk 5 wk	250 mg/kg	104
EGME	mouse B6C3F1	inhal, 6 h/d 9 d	1000 ppm	105
EGME	NZ rabbit	inhal, 6 h/d 5 d/wk, 13 wk	30 ppm	106
EGME	rat, SD	oral, 11 d ^c	100 mg/kg	107
EGME	rat, F344	oral, 5 d	150 mg/kg	108
EGME	rat, F344	inhal 6 h/d 9 d	1000 ppm	105
EGME	rat, SD	inhal, 6 h/d 5 d/wk, 13 wk	300 ppm	110
EGME	rat	inhal, 6 h/d 10 d	800 ppm	111
EGME	rat, alb	inhal, 4 h	625 ppm	112
EGEE	mouse, CD-1	H_2O , 14 wk	1% H ₂ O	113
EGEE	rat, SD	oral, 11 d ^c	500 mg/kg 11 d	107
EGEE	rat	sc, 4 wk oral, 13 wk	400 μl/kg 200 μl/kg	114
EGEE	rat, LE hooded	oral, 5 d	936 mg/kg	115
EGEE	dog	iv, 22 d oral, 13 wk	200 μl/kg	114

^aAbbreviations: NZ, New Zealand; SD, Sprague–Dawley; alb, albino; LE, Long Evans; d, day; h, hour; wk, week; inhal, inhalation; iv, intravenous; sc, subcutaneous.

TESTICULAR TARGET CELLS In rats sacrificed at various times after exposure to EGME (108, 116) and EGEE (116), the meiotic spermatocytes in the pachytene stage were found to be most susceptible to the toxins, whereas cells in earlier (leptotene/zygotene) and later (early spermatid) stages were damaged only after exposure to higher doses or for a longer time period.

These findings were supported by studies in which male rats were mated at various times after exposure to EGME (115, 122, 123). Decreased fertility was observed at the postexposure time period during which the cells exposed at the meiotic spermatocyte stage would have matured to spermatozoa.

Recovery studies suggest that a high dose may cause irreversible damage (107, 110, 115). Effects on fertility or testicular morphology were seen at a postexposure time that would allow a full cycle of spermatogenic maturation, suggesting that the spermatogonia (stem cells) were affected.

^bLowest dose at which treatment-related testicular effects were reported.

^cAnimals sacrificed at various times during the experiments.

GLYCOL ETHER METABOLISM AND ITS RELATIONSHIP TO TESTICULAR TOXICITY The GE that are alkoxyethanols are primarily converted to the corresponding alkoxyacetic acids, which in some cases are conjugated with glycine (124–129). This conversion is thought to be mediated by alcohol dehydrogenase, presumably via the alkoxyacetaldehyde intermediate (127, 129, 130):

R-O-CH₂-CH₂-OH→R-O-CH₂-CHO→R-O-CH₂-COOH [Alkoxyethanol→alkoxyacetaldehyde→alkoxyacetic acid R: CH₃ (EGME), CH₃-CH₂ (EGEE), CH₃-(CH₂)₃· (EGBE)].

Metabolites of EGME and EGEE (and their acetate derivatives) cause testicular toxicity. Methoxyacetic acid (MAA) (107, 130) and ethoxyacetic acid (EAA) (107) produce the same degree of toxicity as the parent compounds, and methoxyaldehyde, the postulated intermediate in the conversion of EGME to MAA, causes similar effects (131). Pyrazole, an alcohol dehydrogenase inhibitor, blocks conversion of EGME to MAA and protects against toxicity (127).

PGME causes none of the toxic effects of EGME, although it differs from EGME only by a single methyl group (105, 111). Unlike the alkoxyethanols, which are primary alcohols, PGME is a secondary alcohol and therefore a poor substrate for alcohol dehydrogenase (132). It is extensively metabolized to CO_2 (128). Metabolism of DPGME, which is also not a testicular toxin, is similar to that of PGME (133).

In primary mixed cultures of Sertoli and germ cells, MAA and EAA caused degeneration of pachytene and dividing spermatozoa (134); these cells are the targets of EGME and EGEE toxicity in vivo (108, 116). EGME and EGEE at much higher concentrations, and *n*-propoxy- and *n*-butoxyacetic acid, metabolites of GE that are not testicular toxins (104, 112, 117–120), had no effect. Furthermore, the ability of the four acids to induce testicular damage in vivo correlated with the in vitro results.

BIOCHEMICAL MECHANISM OF TESTICULAR TOXICITY Exposure of Sertolicells in culture to MAA, but not to EGME itself, decreased production of lactate but had no effect on protein synthesis or cell viability (135). Exposure to EGME decreased in vivo testicular lactate levels (136). These observations are significant because pachytene spermatocytes cannot utilize glucose and are dependent on lactate provided by Sertolicells (137). Other results (134), however, do not support this hypothesis; addition of lactate to mixed Sertoligerm cell cultures did not protect against MAA.

Hematological Effects

ETHYLENE GLYCOL MONOBUTYL ETHER EGBE causes hemolytic anemia, as well as hemoglobinuria; decreased erythrocyte numbers, hemoglobin, and

mean corpuscular volume; and increased mean cell volume, reticulocytes, and nucleated red blood cells; bone marrow and spleen hyperplasia; and increased extramedullary hemopoiesis (117–119). EGPE and its acetate cause similar changes (120).

EGBE may act by increasing the fragility of the erythrocyte membrane, and thus the susceptibility of the cells to hemolysis through its metabolite, butoxyacetic acid (138). However, increased erythrocyte osmotic fragility did not occur after exposure of rats to a concentration sufficient to decrease red cell number.

ETHYLENE GLYCOL MONOMETHYL ETHER In contrast to EGBE, EGME and its acetate derivatives (104) decrease leukocyte numbers (lymphocytes and neutrophils), with less marked effects on erythrocytes (104–106, 111, 119).

EGME depletes the bone marrow of erythroid and myeloid cells (115, 119); abolishes normal extramedullary hemopoiesis (119); damages the endothelial cells of the marrow sinuses (119); and depletes lymphocytes from the thymus (105, 106, 111, 119), spleen (105), and lymph nodes (105). No changes were seen in various immune parameters after treatment with EGME or MAA (139). However, although the dose used produced thymic atrophy, it did not decrease bone marrow cellularity and leukocyte counts, as reported by others. As is the case for testicular toxicity, MAA produces the same hemopoietic toxicity as EGME (139), whereas the closely related PGME has no effect (105, 111).

The observations that EGME causes hematopoietic, lymphoid, and testicular damage might suggest that it acts by a similar mechanism in all of these rapidly dividing tissues. However, other sites of high cell turnover such as the intestinal epithelium and ovarian follicle are unaffected by EGME (105). Also, the most sensitive testicular cells are those undergoing meiosis rather than mitosis (see above).

Developmental Toxicity

A number of studies have shown that EGME, EGEE, and their acetate derivatives, as well as ethylene glycol dimethyl ether (140), can adversely affect embryonic and fetal development (111, 114, 141–148; Table 3). Effects observed include increased incidence of malformations and minor variations, increased embryo mortality, and decreased fetal growth at doses not maternally toxic. Rabbits were more sensitive than rats or mice (143–145). The types of defects caused by EGME are dependent on when in gestation it is given and on how many doses are administered (142).

EGME and EGEE also affect neurochemical and behavioral development in rats (149). Both behavior and neurotransmitter levels changed. Such

Table 3 Selected studies demonstrating the developmental toxicity of glycol ethers

Compound	Species	Exposure	Dose ^a	Reference
EGME	mouse	oral, d.g. 7-14	31.25 mg/kg	141
EGME	mouse	oral, d.g. 7-14 ^c	500 mg/kg 1 dose, 300 mg/kg 3 doses	142 g,
EGME	rabbit	inhal, d.g. 6-18 6 h/d	50 ppm	143
EGME	rat	inhal, d.g. 6-17 6 h/d	100 ppm	111
EGME	rat	inhal, d.g. 7-15 7 h/d	500 ppm	144
EGEE	rabbit	inhal, d.g. 6-18 6 h/d	175 ppm	145
EGEE	rabbit	inhal, d.g. 1-18 7 h/d	160 ppm	146
EGEE	rat	inhal, d.g. 1–19 7 h/d	200 ppm	146
EGEE	rat	oral, d.g. 1-21	50 μl/kg	114
EGEE	rat	inhal, d.g. 6-15 6 h/d	50 ppm	145
EGEE	rat	dermal, d.g. 7-16 4x/d	233 mg	147
EGEEAc	rabbit	inhal, d.g. 6-18 6 h/d	100 ppm	145
EGEEAc	rat	inhal, d.g. 7-15 7 h/d	130 ppm	144
EGEEAc	rat	dermal, d.g. 7-16 4x/d	341 mg	148

^aLowest dose at which developmental toxicity was observed.

changes are also dependent on both the compound and the period of gestation during which exposure occurred.

Evaluations of teratogenic potential of PGME (111, 150), EGBE (144, 148, 151), and EGPE (152) have been negative. These compounds also do not produce testicular toxicity (see above), suggesting that metabolism may also be involved in teratogenicity. Significantly, ethylene glycol itself has recently been found teratogenic in both rats and mice (153, 154). A metabolite (or metabolites) of ethylene glycol, most likely glycolate, are responsible for its systemic toxicity (155, 156); the relationship between ethylene glycol metabolism and its developmental effects has not been examined.

Recent findings suggest a relationship between GE metabolism and teratogenicity. Addition of MAA or EAA blocked growth and development of

bAbbreviation: d.g., days of gestation.

^cMultiple or single dose given during the period.

rat embryos in culture (157). Propoxyacetic, butoxyacetic, and methoxypropionic acids, the analogous derivatives of the nonteratogenic GE (158), as well as EGME itself (157), showed much weaker toxicity in this system.

MAA is teratogenic in rats (159). EGME, MAA, and dimethoxyethyl phthalate were equally potent teratogens when administered to rats on day 12 of gestation (160). Metabolites of dimethoxyethyl phthalate include phthalate and EGME (161). The three compounds induced similar types of malformations, including several unusual defects not induced by other teratogens.

MAA accumulates in the fetus after injection into pregnant rats (162), and radioactivity from 2-methoxy[1,2-¹⁴C]ethanol injected into pregnant mice is incorporated into fetal macromolecules (163). The alcohol dehydrogenase inhibitor, 4-methyl pyrazole (160), or ethanol (164), which would be expected to compete with EGME for metabolism, protects the fetus from teratogenic effects. However, fetal accumulation of MAA is not decreased by ethanol. This finding suggests that subsequent metabolism of MAA may be required for toxicity and that this metabolism may be blocked by ethanol (165). In support of this idea, coadministration of formate or acetate decreased the incidence of paw malformations from EGME or MAA. Thus, competition of metabolites of MAA with endogenous carboxylic acids may contribute to the developmental toxicity of EGME (166).

MECHANISM OF CARBON TETRACHLORIDE-INDUCED HEPATOTOXICITY

Among toxic effects of CCl₄ in liver is a decrease in xenobiotic metabolism. Recent studies suggest that CCl₄ acts as a suicide substrate for cytochrome P-450 (167). The phenobarbital-induced form of the cytochrome P-450 is the most susceptible to attack. The trichloromethyl free radical, which results from the homolytic cleavage of CCl₄, may bind either at the heme group of cytochrome P-450 or at the active site of the enzyme near the heme group, thereby leading to inactivation (168, 169). Lipid peroxidation, which has been thought for many years to play a seminal role in CCl₄ toxicity, may relate directly to the decrease in cytochrome P-450 (170). In addition to the well-established route of lipid peroxidation (171), metabolism of [14C]CCl₄ in rat liver microsomes was found to produce trichloromethyl free radicals, which were covalently bound to phospholipids (172). Isolation of phosphatidylcholine from the phospholipid fraction and incubation with phospholipase A2 demonstrated that about half of the phospholipid had been rendered resistant to hydrolysis. Thus, it has been argued that whereas direct reaction of the radical with the heme moiety of cytochrome P-450 leads to destruction of the mixed-function oxidase, lipid peroxidation is more closely related to loss of microsomal enzymes such as glucose 6-phosphatase, UDP-glucuronyl-transferase, nucleoside diphosphatase, and perhaps other enzymes (173).

Hypoxia appears to potentiate CCl₄-induced hepatotoxicity (174). The production of ethane and pentane (thought to be indicators of lipid peroxidation in vivo, in an oxygen deficient atmosphere) from rats given CCl₄ reached a plateau at normal oxygen pressure (175) and reached a higher plateau at reduced oxygen levels. The initial phase of metabolism is thought to plateau because reactive metabolites of CCl₄ destroy cytochrome P-450. The reaction is more rapid anerobically. The metabolism of CCl₄ appeared to be accompanied by the destruction of a single, rather than a large number, of fatty acids. When oxygen tension was measured as a function of oxygen consumption and lipid peroxidation, as determined by malondialdehyde formation in rat liver microsomes during CCl₄ metabolism, both parameters were greatest in the range of 1–10 mm Hg (0.1–1.3 kPa) but were less at 80 mmHg (10.7 kPa) (173, 176).

Rechnagel & Glende (177) postulated that CCl₄ is metabolized via a homolytic splitting of the carbon-chlorine bond, which probably occurs anerobically following the reaction of CCl₄ with cytochrome P-450 (Fe²⁺) (Figure 2) (178). The result is the formation of a complex composed of cytochrome P-450 (Fe³⁺) and trichloromethyl free radical. This structure is analogous to the cytochrome P-450-oxy complex, which is probably involved in oxidative reactions. The trichloromethyl free radical may either be released from its complex in a manner similar to that by which the superoxide anion radical is released from cytochrome P-450 or it may undergo a one-electron reduction to yield [Fe³⁺—|CCl₃↔Fe²⁺—·CCl₃]. The trichloromethyl free radical can abstract a proton and form chloroform or, upon further reduction, yield a ferrodichlorocarbene complex (179), which can release HCl and CO in the presence of water. The mechanism of degradation of the carbene is probably analogous to that proposed by Kubic & Anders (180) for the thiol-stimulated conversion of dihalomethanes to CO (181). The residual cytochrome P-450 (Fe²⁺) is free to react with CC1₄ again.

Slater (182) suggested that the trichloromethyl free radical may be less reactive than previously thought, but that it could react with oxygen to form a trichloromethyl peroxy free radical, which is a more reactive species (183, 184). "Unequivocal evidence" that trichloromethyl free radicals are formed during the metabolism of CCl₄ was provided in studies employing isolated rat hepatocytes and spin traps to identify the free radical species (185). Unsaturated lipid free radicals were also detected. There was no evidence that the trichloromethyl peroxy free radical was formed until the recent report of Connor et al (186). These investigators trapped a carbon dioxide anion radical resulting from CCl₄ metabolism in perfused liver using the spin trap *N*-t-butylnitrone. The authors concluded that the carbon dioxide anion radical

Figure 2 Formation of reactive metabolites from CCl₄.

arose from the trichloromethyl peroxy free radical. The trichloromethyl peroxy free radical is thought to be a precursor to phosgene, which may also play a role in CCl₄-induced hepatotoxicity (187).

McCay et al (188) also used spin-trapping techniques to identify the trichloromethyl free radical in rat liver microsomes. They suggested that free radicals may be formed in a series of events initiated when a free radical reacts with a lipid moiety (LH; Figure 2). For example, during the homolytic cleavage of the C-Cl bond, radicals are formed that may react with lipid to yield L·, which can in turn react with O₂ to form the LOO· radical. LOOH may take either of two routes. Its action depends on the oxidation state of iron and results in the production of a highly complex series of reaction products (189). A similar series of reactions probably occurs in the liver.

The trichloromethyl radical can react directly with lipids or can be converted to the trichloromethyl peroxy radical; formation of the peroxy radical appears to depend on the availability of oxygen. The following series of events probably takes place in the hepatocyte: CC1₄ is reductively cleaved by cytochrome P-450 (190). This reduction yields trichloromethyl free radical in a region vicinal to unsaturated phospholipids. Thus, some radicals can add directly to double bonds to yield dienyl radicals. Others react with oxygen to form trichloromethyl peroxy radicals, which may also act in the initiation of lipid peroxidation. The lipid radicals resulting from the free-radical attack may react with oxygen as described above to initiate the process of lipid peroxidation, as Kappus (171) has suggested. Thus, oxidation of omega-3

unsaturated fatty acids yields ethane, whereas oxidation of omcga-6 unsaturated fatty acids yields *n*-pentane.

Disturbances in Calcium Ion Homeostasis as a Final Common Pathway in Some Forms of Solvent Hepatotoxicity

Rechnagel (191), in attempting to relate CCl₄-induced lipid peroxidation to hepatic cell death and necrosis, recognized that lipid peroxidation occurred primarily in the region surrounding cytochrome P-450. He first attempted to invoke a useful concept, which he termed a "toxicological second messenger," to link events in the endoplasmic reticulum with general cell injury. The candidates he thought most likely to possess the required properties were 4-hydroxyalkenals, which are highly toxic products of lipid peroxidation. However, despite the attractiveness of this hypothesis, he recognized that 4-hydroxyalkenals react rapidly with cellular constituents in close proximity to their points of origin and would be unlikely to act as messengers to other loci in the cell. An alternative hypothesis developed by Rechnagel (191) and by Orrenius et al (192) suggests that alterations in Ca²⁺ homeostasis resulting from cell injury by CC14 or bromobenzene lead to cell death. Farber et al (193) had previously suggested that cell injury due to ischemia could be prevented if large increases in free intracellular Ca²⁺ could be prevented. They had argued that toxic liver cell death was associated with susceptibility of the cells to a large influx of extracellular Ca²⁺ (194). However, it was demonstrated that CCl₄, bromobenzene, and ethyl methane sulfonate were more toxic to liver cells in the absence of extracellular Ca2+ than in its presence. This finding suggested that redistribution of intracellular Ca2+ might be critical for cell death. CC14 inhibits the ability of microsomes to sequester Ca²⁺ but does not prevent the influx of extracellular Ca²⁺ either in vivo (195) or in vitro (196). Rechnagel (191, 197) argued that release of Ca²⁺ into the cytosol could result in a number of regulatory alterations. Such alterations could lead to triglyceride accumulation, a prominent feature of CCl₄-induced hepatotoxicity.

Bromobenzene caused a depletion of cellular GSH concomitant with the appearance of blebs on the surface of isolated hepatocytes (198). The blebs were postulated to be related to changes in membrane permeability associated with bromobenzene toxicity. Menadione, which also forms surface blebs, served as a model for hepatotoxicity induced by bromobenzene and other hepatotoxins. In this study Ca²⁺ accumulated in isolated perfused liver cells (199). Blebbing was also induced by the Ca²⁺ ionophore, A23187 (200), in a Ca²⁺-free medium, suggesting redistribution of intracellular Ca²⁺. Bellomo et al (201) also demonstrated impairment of calcium sequestration in mitochondria and the mobilization and loss of Ca²⁺ from both mitochondria

and extramitochondrial spaces (199). Glutathione can prevent these effects. Orrenius et al (192), in agreement with Rechnagel (191), argued that these effects will cause the release of Ca²⁺ into the cytosol. They suggested that membrane-bound Ca2+-ATPase, a sulfhydryl enzyme, is also inactivated by reactive intermediates and can no longer correct the imbalance of Ca²⁺ by promoting its secretion. The exact mechanism by which increased intracellular Ca²⁺ causes cell death is not yet completely understood. Further studies are needed to determine whether Ca²⁺ accumulation resulting from the cellular damage that reactive intermediates of solvent metabolism induce represents a final common pathway for cell death.

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